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(54) **High surface area ceria.**

(57) A thermally stable, high surface area ceria composition is obtained by intimately adding a ceria stabilizer of one or more of La, Nd or Y to the ceria in an amount of from about 5 to 25 mole % of the ceria. One way to make the product is to mix the stabilizer with a ceria precursor such as cerium nitrate, evaporate the mixture and then calcine the salts to form the stabilized ceria. Another way to make the product is to form an aqueous solution of precursor salts which are precipitated with a base to form hydrous cerous and stabilizer hydroxides or their carbonates. This mixture can be washed and then calcined to form the stabilized ceria. After heating in air for 4 hours at 900 °C the stabilized ceria retains a surface area of greater than 20 m² /g.

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Background of the Invention

1. Field of the Invention

5 This invention relates to a ceria preparation route which produces thermally stable, high surface area ceria. The ceria can be incorporated into a catalytic washcoat.

2. Description of the Previously Published Art

10 Automotive three-way catalytic converters are employed to convert pollutants produced by internal combustion engines into harmless emissions. The catalyst must be highly active for converting CO, NO, and unburned hydrocarbons into nitrogen, carbon dioxide and water over a wide range of exhaust compositions. As a result, a noble metal catalyst which is usually platinum, rhodium, and/or palladium must be employed. To utilize the noble metal most efficiently, they are dispersed on a high surface area support
 15 such as gamma alumina (typically having a surface area greater than 150 m²/g). The high surface area support also provides a surface for the deposition of poisons such as lead, zinc, and phosphorus which are usually present in the exhaust gas. Since the catalyst may reach temperatures as high as 1000 °C, the support must be thermally stable and resistant to sintering.

The composition of the engine exhaust gas oscillates between net oxidizing (lean) for higher than
 20 stoichiometric air/fuel ratios and net reducing (rich) for lower than stoichiometric air/fuel ratios. However, the catalyst operates most efficiently in a stoichiometric exhaust gas. Ceria is commonly added to three-way catalysts to promote the activity of the noble metals under these transient conditions. Ceria serves as an oxygen storage component, since Ce³⁺ can be oxidized to Ce⁴⁺ under lean conditions; and, similarly, Ce⁴⁺ can be reduced to Ce³⁺ under rich conditions. In this manner, ceria tends to dampen out oscillations in the
 25 exhaust gas stoichiometry.

Ceria may be incorporated into a three-way catalyst washcoat either by impregnation of the high surface area alumina support with a cerium salt or by physically mixing bulk cerium oxide with the other catalyst components. The addition of bulk ceria to the catalyst may be preferable since interactions between impregnated ceria and the alumina support decrease the oxygen storage capacity of the ceria.
 30 However, at the extreme temperatures often encountered in catalytic converters (in excess of 1000 °C), ceria rapidly sinters to surface areas of less than 5 m²/g. Therefore, the incorporation of a thermally stable, high surface area ceria would prove beneficial to three-way catalyst performance.

U. S. Patent No. 3,830,756 discusses the importance of using a high surface area ceria support for palladium and platinum oxidation catalysts. Ceria with a purity of 99% was prepared by precipitation of
 35 cerium nitrate with ammonium hydroxide at a pH of 9. While the ceria had a high surface area of 218 m²/g after calcination at 427 °C for 4 hours in air, the surface area dropped to 4 m²/g following a 4 hour calcination at 982 °C.

U. S. Patent No. 4,661,330 discloses a method for preparing high surface area, high purity ceria. Ammonium ceric nitrate is refluxed for 24 hours with ammonium sulfate to obtain a hydrous ceria powder.
 40 Following calcination at 538 °C in air, this material maintains a surface area of 150 m²/g. However, the stability of this material at higher calcination temperatures was not investigated.

U. S. Patent No. 4,714,694 presents a method for stabilizing the surface area of ceria. A high surface area ceria, calcined at a low temperature, is impregnated with an aluminum salt and calcined. Ceria prepared by this method containing 8 mole % aluminum has a surface area of only 10.5 m²/g following
 45 calcination at 1000 °C in air.

U. S. Patent 4,859,432 teaches a method for preparing a morphologically improved ceria having high surface area. Ceria is produced by reacting a cerium salt with a strong base in the presence of carboxylate ions followed by calcination. After calcination at 600 °C, ceria with a surface area of 117 m²/g was obtained. Higher calcination temperatures were not investigated.

50 In contrast to these poor products of the prior art, the present invention identifies additives and preparation methods for preparing a ceria which maintains a surface area in excess of 20 m²/g as measured by the Brunauer-Emmett-Teller (BET) method following a calcination at 980 °C for 4 hours in air.

3. Objects of the Invention

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It is an object of this invention to produce a thermally stable, high surface area ceria.

It is further object of this invention to produce a high surface area ceria which is unusually resistant to sintering.

It is further object of this invention to identify additives which increase the surface area of ceria following thermal treatments.

It is further object of this invention to develop a method by which high surface area ceria may be prepared.

5 These and further objects will become apparent as the description of the invention proceeds.

Summary of the Invention

10 Ceria is produced which is unusually resistant to sintering. The addition of rare earths to the ceria during its preparation stabilizes the high surface area ceria which is obtained. For the purposes of this invention, a thermally stable, high surface area ceria is defined as a material which maintains a surface area of over 20 m²/g as measured by the Brunauer-Emmett-Teller (BET) method following calcination at 980 °C in air for at least 4 hours.

15 In a preferred procedure, a ceria precursor such as cerium nitrate is mixed with a stabilizer precursor compound such as lanthanum, neodymium, or yttrium nitrate and the mixture is dissolved in water.

The precursor mixtures range from about 5 to 25 mole % and preferably from about 5 to 20 mole % stabilizer (lanthanum, neodymium, or yttrium). The precursor solutions are either evaporated to give a mixed rare earth salt or precipitated at various temperatures and pH with various precipitating agents to precipitate the mixture as a hydroxide or a carbonate. The precipitate may be filtered and washed free of excess salts 20 using water, ammonia in water, methanol, or other solvents. The preparations are then calcined to produce a stabilized ceria.

Description of the Preferred Embodiments

25 Addition of a stabilizing component to ceria produces a thermally stable, high surface area ceria. The manner in which the stabilizer is incorporated influences the properties of ceria.

Ceria can be prepared by calcination of precursor compounds including cerous acetate, bromide, chloride, iodide, nitrate, or sulfate; ceric sulfate; ammonium cerous sulfate or nitrate; or ammonium ceric sulfate or nitrate. Alternatively, ceria may be prepared by precipitation of hydrous ceric or cerous oxide 30 from aqueous solutions of these precursors with ammonium hydroxide, caustic soda, or other alkali metal hydroxides. The precipitate is then filtered, washed with water, methanol, or another solvent to remove the excess salts, dried, and calcined in air at a temperature in excess of 500 °C to produce ceria. Alternatively, cerous or ceric carbonate may be precipitated by addition of ammonium carbonate, sodium carbonate, or other alkali metal carbonates to aqueous solutions of these precursors. The precipitate may then be treated 35 as described above to produce ceria.

High surface area ceria may be produced by incorporating 5 to 25 mole %, preferably 5 to 20 mole % and most preferably about 10 mole %, of lanthanum, yttrium, or neodymium. The preferred composition is to incorporate lanthanum into the ceria. The lanthanum, yttrium, or neodymium precursors used include the acetate, bromate, bromide, chloride, iodide, nitrate, and sulfate salts, preferably the nitrate. The stabilized 40 ceria may be prepared by dissolving the cerium precursor, preferably nitrate, in water with the stabilizer precursor, preferably the nitrate, in the appropriate ratio to give the desired mole % of precursor. The initial concentration of nitrates in the solution is preferably 500 g/l. The aqueous mixture of the cerium salt and the salt of the stabilizer can then be evaporated to produce an intimate mixture of the precursor salts which is then calcined in air at a temperature above 500 °C to produce a stabilized ceria.

45 Alternatively, using the preferred procedure, the aqueous mixture of the precursor salts may be combined with aqueous solutions of ammonium hydroxide, caustic soda, or other alkali metal hydroxides to produce an intimate mixture of the hydrous cerous and stabilizer hydroxides. For this procedure, the initial concentration of the precursor salts may be 10-500 g/l, preferably 100 g/l. The concentration of the precipitating agent may be 1-25 g/100 g of solution and preferably 5 g/100 g of solution. The precipitate is 50 then filtered, washed with water, methanol, or another solvent to remove the excess salts, dried in air at 120 °C for 8 hours, and calcined in air at a temperature in excess of 500 °C to produce stabilized ceria. Alternatively, the precursors may be precipitated as carbonates by the addition of ammonium carbonate, sodium carbonate, or the carbonates of other alkali metals to the aqueous solution of these precursors. The precipitate may then be washed, dried, and calcined as described above to produce a stabilized ceria.

55 The precipitation described above may be performed at the ambient temperature, or, preferably, at an elevated temperature, preferably 70-85 °C. The precipitation may be performed by adding the solution containing the precipitating agent to a stirred volume of the precursor solution, by adding the precursor solution to a stirred volume of the solution containing the precipitating agent, or, preferably, by simulta-

neously adding the precursor solution and the solution containing the precipitating agent to a stirred volume of water while maintaining a constant pH between 7 and 10, and preferably 9.

Uniform incorporation of the ceria stabilizer into the ceria precursor to form an intimate mixture prior to calcining the precursor to form the ceria is necessary in order to provide a high surface area ceria. As a result, these preparation methods involving dissolution of precursor compounds so that it can be intimately mixed with the ceria precursor are preferred. Other methods such as impregnation of an already formed crystalline ceria with a stabilizer or mixing and firing of cerium oxides and stabilizer oxides yield inferior materials.

Having described the basic aspects of our invention, the following examples are given to illustrate specific embodiments thereof.

Example 1

This example illustrates the preparation of stabilized ceria by evaporation.

Stabilized ceria containing 10 mole % La and 90 mole % Ce (sample C in Table I) is prepared by adding 45.657 g of cerium nitrate and 5.059 g of lanthanum nitrate to 80 g of water, and the solution is heated until all of the solid has dissolved. Additional water is then added to bring the total volume of the solution to 101 ml (concentration of salts is 500 g/l). The solution is then evaporated to dryness in air, and the resulting solid is calcined in air at 538 °C for 1 hour. Using this procedure, cerias containing 0, 5, and 20 mole % La are also prepared (samples A, B, and D in Table I, respectively). The BET surface areas of these materials following calcination in air at 538 °C, 760 °C, and 980 °C for 4 hours are given in Table I.

Table I

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Sample	La Content Mole %	BET Surface Area (m ² /g) ^a		
		538°C	760°C	980°C
A	0	86	30	1.4
B	5	59	33	15
C	10	67	35	22
D	20	64	46	20

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(a) After calcination at listed temperature for 4 hours.

The data in Table I show that the incorporation of La into ceria greatly improves the thermal stability of ceria. Following calcination at 980 °C for 4 hours, ceria containing 10 mole % La has a sixteen-fold higher surface area than unstabilized ceria. In addition, the optimum composition appears to be 10 mole % La with 90 mole % Ce.

Example 2

This example illustrates the preparation of stabilized ceria by precipitation.

Stabilized ceria containing 10 mole % La and 90 mole % Ce (sample F in Table II) is prepared by adding 45.657 g of cerium nitrate and 5.509 g of lanthanum nitrate to 400 g of water, and the solution is heated to 70 °C. Additional water is added to bring the total volume to 507 ml (corresponding to a salt concentration of 100 g/l). A solution of 5% NH₃ in water is prepared at 25 °C. These two solutions are added simultaneously to 500 ml of water that is kept at a constant temperature of 70 °C. The nitrate solution is added at a rate of 50 ml/min, and the relative rates of addition are controlled to keep the pH of the solution constant at 9. The resulting precipitate is then filtered, washed with 2 l of deionized water, and dried at 120 °C in air for 8 hours. The resulting solid is calcined at 538 °C in air for 1 hour. Using this

precipitation procedure, a sample containing no La is also prepared (sample E in Table II). A similar procedure, identical to that described with the exception that a precipitation temperature of 85°C is employed, is used to prepare samples containing 0 and 10 mole % La (samples G and H in Table II, respectively). The BET surface areas of these samples following calcination at 980°C in air for 4 hours are given in Table II.

TABLE II

Sample	La Content	T precipitation	BET Surface Area
	Mole %	(°C)	(m ² /g) 980°F
E	0	70	5
F	10	70	25
G	0	85	5
H	10	85	34

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The data in Table II demonstrate a five to seven-fold increase in surface area following calcination at 980°C by the incorporation of La into a ceria during the precipitation process. Furthermore, conducting the precipitation at 85°C instead of 70°C results in further gains in surface area. A comparison of samples G and H in Table II with samples A and C in Table I demonstrates that precipitation is preferred over evaporation as a method for preparing high surface area ceria.

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Example 3

This example illustrates the preparation of ceria stabilized by Y and Nd.

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Stabilized ceria containing 10 mole % Y and 90 mole % Ce (sample I in Table III) is prepared by adding 47.031 g of cerium nitrate and 4.609 g of yttrium nitrate to 400 g of water, and the solution is heated to 85°C. Additional water is added to bring the total volume to 507 ml (corresponding to a salt concentration of 100 g/l). A solution of 5% NH₃ in water is prepared at 25°C. These two solutions are added simultaneously to 500 ml of water that is kept at a constant temperature of 85°C. The nitrate solution is added at a rate of 50 ml/min, and the relative rates of addition are controlled to keep the pH of the solution constant at 9. The resulting precipitate is then filtered, washed with 2 l of deionized water, and dried at 120°C in air for 8 hours. The resulting solid is calcined at 538°C in air for 1 hour. Using this precipitation procedure, a sample containing 10 mole % Nd and 90 mole % Ce is also prepared (sample J in Table III). The BET surface areas of these samples following calcination at 980°C in air for 4 hours are given in Table III.

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TABLE III

Sample	Additive (mole%)	BET Surface Area (m ² /g)
		980°C
I	10% Y	26
J	10% Nd	26

Comparison of the results in Table III with samples G and H in Table II demonstrate that incorporation of Y and Nd into ceria greatly improves the thermal stability of ceria. However, La is more effective than Y or Nd at producing a high surface area ceria.

Example 4

This example illustrates the preparation of stabilized ceria with a methanol wash.

Stabilized ceria containing 10 mole % La and 90 Mole % Ce (sample K in Table IV) is prepared by adding 45.657 g of cerium nitrate and 5.059 g of lanthanum nitrate to 400 g of water, and the solution is heated to 85°C. Additional water is added to bring the total volume to 507 ml (corresponding to a salt concentration of 100 g/l). A solution of 5% NH₃ in water is prepared at 25°C. These two solutions are added simultaneously to 500 ml of water that is kept at a constant temperature of 85°C. The nitrate solution is added at a rate of 50 ml/min, and the relative rates of addition are controlled to keep the pH of the solution constant at 9. The resulting precipitate is then filtered, washed with 2 l of deionized water, and washed with 500 ml of methanol, and dried at 120°C in air for 8 hours. The resulting solid is calcined at 538°C in air for 1 hour. Using this procedure, a sample containing no La is also prepared (sample L in Table IV). The BET surface areas of these samples following calcination at 980°C in air for 4 hours are given in Table IV.

TABLE IV

Sample	La Content (mole%)	BET Surface Area (m ² /g)
		980°C
K	0	6
L	10	47

Comparison of the results in Table IV with sample H in Table II demonstrates that washing the precipitated stabilized ceria with methanol following the water wash results in further improvements in thermal stability.

Example 5

This example provides a comparison with production of stabilized ceria prepared according to a prior art method.

Ceria is prepared according to the procedure described in U. S. Patent No. 4,661,330 (sample M in Table V). Into 1090 ml of deionized water 37.4 g of ammonium sulfate is dissolved. Into 400 ml of deionized water 426.386 g of ammonium ceric nitrate is dissolved, and the total volume is brought to 610 ml. The ceric solution is added to the ammonium sulfate solution and stirred continuously while heated to boiling. The solution is refluxed for 24 hours. The resulting precipitate is filtered, washed with 2 l of 90 °C deionized water, and dried for 8 hours in air at 120 °C. The resulting solid is calcined in air at 538 °C for 1 hour.

To see the effect of using the preferred stabilizer, La, according to the present invention, a ceria stabilized composition containing 10 mole % La and 90 mole % Ce was prepared by an identical procedure, except that the ceric solution was prepared to contain 383.747 g of ammonium ceric nitrate and 33.678 g of lanthanum nitrate (sample N in Table V) so that the total moles of rare earth would be the same. The BET surface areas of these samples following calcination at 980 °C in air for 4 hours are given in Table V.

TABLE V

Sample	La Content (mole%)	BET Surface Area (m ² /g)
		980 °C
M (prior art)	0	5
N (modified prior art)	10	13

Although, as reported in U. S. Patent 4,661,330, this method of ceria preparation may produce a high surface area ceria for low calcination temperatures (500 °C or less), Table V shows that the ceria does not maintain a high surface area for a calcination temperature of 980 °C. Table V also shows that incorporating La (which is a stabilizer preferred in the present patent) is effective at increasing the surface area of ceria prepared according to U.S. Patent No. 4,661,330. However, the superior stabilized ceria according to the present invention having a surface area of greater than 20 m²/g is not obtained. Thus to prepare a thermally stable high surface area ceria, the precipitation method taught by the present invention is clearly preferred.

Example 6

This example provides a comparison with a stabilized ceria prepared according to another prior art method.

Ceria stabilized with Al is prepared according to the procedures described in U.S. Patent No. 4,714,694 (sample O in Table VI). Ceria with a purity of 99% and a BET surface area of 120 m²/g is impregnated with aluminum nitrate to provide a loading of 10 mole % Al with 90 mole % Ce. The sample is dried at 120 °C in air for 8 hours and calcined at 538 °C in air for 1 hour. Ceria stabilized with La (a stabilizer according to the present invention) containing 10 mole % La and 90 mole % Ce is prepared by an identical procedure, except that ceria is impregnated with La nitrate instead of Al nitrate (sample P in Table VI). The BET surface areas of these samples following calcination at 980 °C in air for 4 hours are given in Table VI.

TABLE VI

Sample	Additive (mole%)	BET Surface Area (m ² /g)
		980°C
O	10% Al	8
P	10% La	13

The results in Table VI show that using La (which is a stabilizer according to the present invention) is preferable to using Al as a stabilizer as taught by U.S. Patent No. 4,714,694. However, the superior stabilized ceria according to the present invention having a surface area greater than 20 m²/g is not obtained. Thus the precipitation procedure for stabilizer incorporation as taught in the present invention is also preferred to the post impregnation of ceria procedure of U. S. Patent 4,714,694.

Example 7

This example provides a comparison with a stabilized ceria prepared according to a prior art method. Ceria was prepared according to the procedure described in U. S. Patent No. 4,859,432 (sample Q in Table VII). 1.25 liter of a 5 normal sodium hydroxide solution was prepared and stirred at 500 rpm at 20 °C. Over 15 minutes, 397 g of solid cerium acetate was added. The temperature of the mixture was raised to 80 °C for 4 hours, and then allowed to cool to room temperature. To this mixture, 95.4 milliliter of a 9.8 molar hydrogen peroxide solution was added. The temperature of the mixture was then raised to 70 °C for 1 hour and then cooled to room temperature. The precipitate was filtered and washed twice by suspending the precipitate in water at 100 g of precipitate per liter of water, and a third was was conducted using a dilute nitric acid solution at pH 5. The solid was then dried at 120 °C for 8 hours in air and calcined in air at 538 °C for 1 hour.

A sample containing the preferred stabilizer, La, according to the present invention, was prepared by an identical procedure, except that 47.700 g of lanthanum acetate were added simultaneously with the cerium acetate to produce a sample containing 10 mole % La and 90 mole % Ce (sample R in Table VII). The BET surface areas of these samples following calcination at 980 °C in air for 4 hours are given in Table VII.

TABLE VII

Sample	La Content (mole %)	BET Surface Area (m ² /g)
		980°C
Q (prior art)	0	1
R (modified prior art)	10	7

The results in Table VII demonstrate that the procedure reported in U. S. Patent No. 4,859,432 does not produce a thermally stable ceria for a calcination temperature of 980 °C, even if La is used to stabilize the ceria. Thus, the precipitation method taught by the present invention is preferred.

It is understood that the foregoing detailed description is given merely by way of illustration and that many variations may be made therein without departing from the spirit of this invention.

Claims

1. A thermally stable, high surface area ceria composition comprising ceria and from about 5 to 25 mole % based on the moles of ceria of a ceria stabilizer selected from the group consisting of La, Nd, Y and mixtures thereof.
2. A ceria composition according to Claim 1, wherein after heating the composition to a temperature of 980 °C for four hours in air, the composition retains a BET surface area of greater than 20 m² /g.
3. A ceria composition according to Claim 1, wherein the ceria stabilizer is present from about 5 to 20 mole % based on the moles of ceria.
4. A ceria composition according to Claim 3, wherein the ceria stabilizer is present in about 10% mole based on the moles of ceria.
5. A ceria composition according to Claim 1, wherein the ceria stabilizer is La.
6. A ceria composition according to Claim 1, wherein the ceria stabilizer is Nd.
7. A ceria composition according to Claim 1, wherein the ceria stabilizer is Y.
8. A method of producing thermally stable, high surface area ceria composition comprising the steps:
 - (a) mixing a ceria precursor with from 5 to 25 mole % of a ceria stabilizer selected from the group consisting of La, Nd, Y and mixtures thereof;
 - (b) forming an intimate mixture of the ceria precursor and the ceria stabilizer by either evaporation of the mixture of step (a) or precipitation of the mixture of step (a) as a hydroxide or a carbonate; and
 - (c) calcining the resulting intimate mixture to obtain stabilized ceria.
9. A method according to Claim 8, wherein the ceria stabilizer is added in step (a) in the form of a ceria stabilizer precursor compound.
10. A method according to Claim 9, wherein the ceria stabilizer precursor is a salt selected from the group consisting of acetate, bromate, bromide, chloride, iodide, nitrate, sulfate, and mixtures thereof.
11. A method according to Claim 8, wherein in step (b) the intimate mixture is formed by evaporation of the mixture.
12. A method according to Claim 8, wherein in step (b) the intimate mixture is formed by precipitation of the ceria and ceria stabilizer from precursor compounds as hydroxides or carbonates.
13. A method according to Claim 12, wherein the precipitation as hydroxides is performed by adding a solution of ammonium hydroxide, caustic soda, or other alkali metal hydroxides.
14. A method according to Claim 12, wherein the precipitation as carbonate is performed by adding a solution of ammonium carbonate, sodium carbonate, or other alkali metal carbonates.
15. A method according to Claim 12, wherein the precipitate is filtered and washed free of other salts.
16. A method according to Claim 15, wherein the washing is done with water, ammonia in water, methanol or other solvents.
17. A method according to Claim 8, wherein the ceria stabilizer is present from about 5 to 20 mole % based on the moles of ceria.
18. A method according to Claim 17, wherein the ceria stabilizer is present in about 10 mole % based on the moles of ceria.
19. A method according to Claim 8, wherein the ceria precursor is selected from the group consisting of

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cerous acetate, cerous bromide, cerous chloride, cerous iodide, cerous nitrate, cerous sulfate, ceric sulfate, ammonium cerous sulfate, ammonium cerous nitrate, ammonium ceric sulfate, ammonium ceric nitrate, and mixtures thereof.

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EUROPEAN SEARCH REPORT

Application Number

EP 91 10 1969

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.5)
X	EP-A-0 207 857 (RHONE-POULENC) * Claims 1-5,10-14; examples 3,1 * - - -	1-5,8-11, 19	B 01 J 23/10 C 01 F 17/00
X	FR-A-2 201 277 (MITSUBISHI CHEMICAL) * Pages 7-8 * - - -	1-8,12, 13,15,19	
X	US-A-4 008 254 (D.E. GROSS) * Example 5 * - - -	1-8,17,18	
A	US-A-3 892 836 (W.A. COMPTON) - - - - -		
			TECHNICAL FIELDS SEARCHED (Int. Cl.5)
			B 01 J C 01 F
The present search report has been drawn up for all claims			
Place of search The Hague		Date of completion of search 29 May 91	Examiner THION M.A.
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